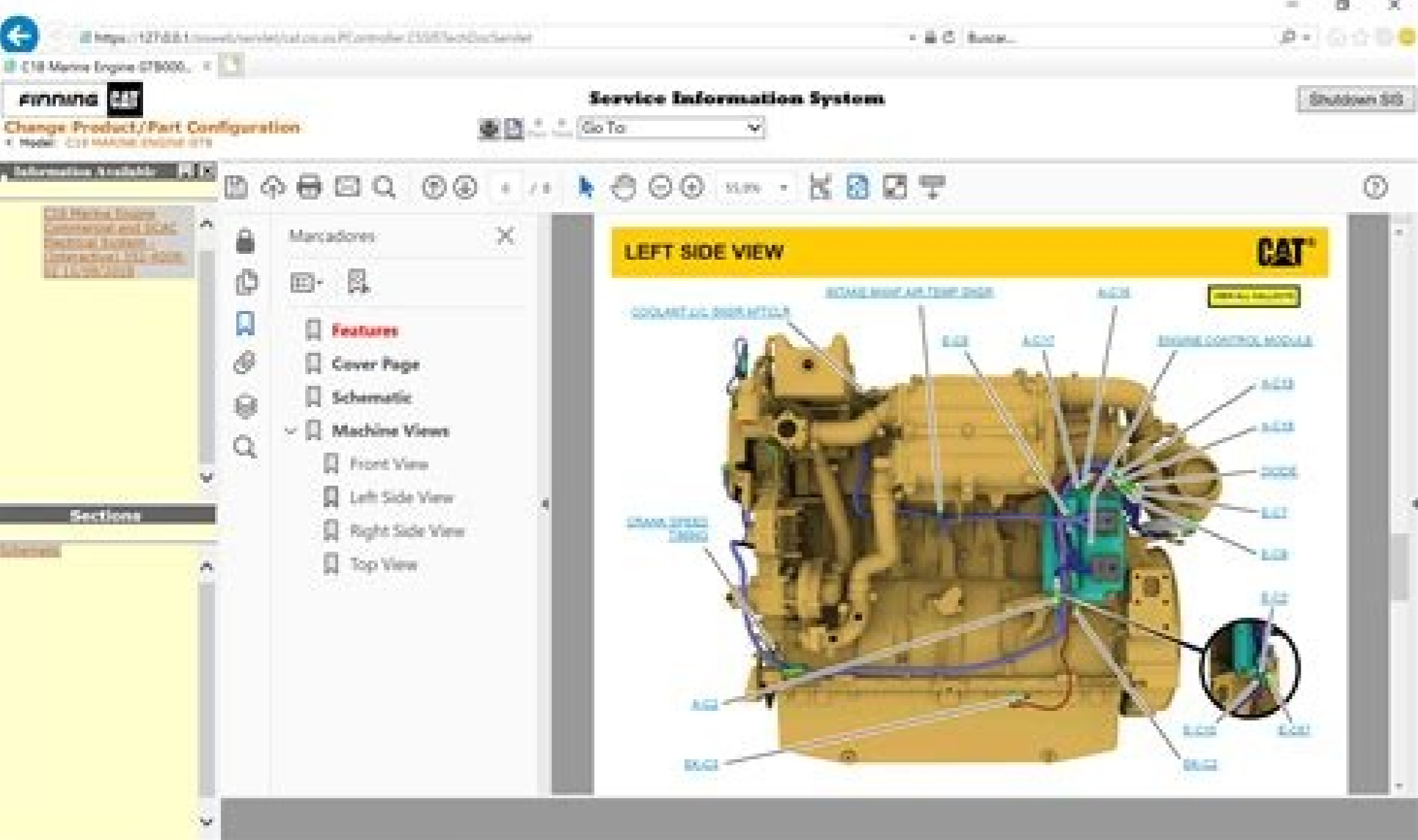


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Now we will build the configuration of the ground state electrons and the orbital diagram for a selection of atoms in the first and second periods of the periodic table. The orbital diagrams are pictorial representations of the electron configuration, showing individual orbitals and the arrangement of electron coupling. We start with a single atom of hydrogen (atomic number 1), which consists of a proton and an electron. Referring to the figure \ (PageIndex{3}) or \ (PageIndex{4}), we expect to find the electron in the 1s orbital. By convention, the $(m_s = +\frac{1}{2})$ value is usually filled first. The electron configuration and the orbital diagram are: Following hydrogen is the noble gas helium, which has an atomic number of 2. The helium atom contains two protons and two electrons. The first electron has the same four quantum numbers of the hydrogen atom electron ($n = 1, l = 0, m_l = 0, m_s = +\frac{1}{2}$). The second electron also enters orbital 1 and fills that orbital. The second electron has the same quantum number n, l and m_l , but must have the opposite rotation number, $(m_s = -\frac{1}{2})$. This is in accordance with the principle of exclusion Pauli: No two electrons in the same atom can have the same set of four quantum numbers. For orbital diagrams, this means that two arrows go in each box (representing two electrons in each orbital) and the arrows must point in opposite directions (representing matching turns). The configuration of the electron and the orbital diagram of the helium are: The $n = 1$ shell is completely filled in an atom of helium. The next atom is alkali metal lithium with an atomic number of 3. The first two lithium electrons fill the 1 orbitals and have the same sets of four quantum numbers of the two helium electrons. The remaining electron mustThe orbital of the next lowest energy, the 2s orbital (figure \ (PageIndex{3}) or \ (PageIndex{4})). So, the p_2 orbital ligned ou ni inorttele id aippoc anu ah j6 ocimota oremun(onegisso'L .Itaipocca non irig onnah inorttele ert itseuQ ,dnuH id aloger al noc .Atimrofnoc ni ,p2 elatibro ert led onucaic ni enorttele nu ah e S2 e S1 ellehsottos el eipmeir j7 ocimota oremun(otoza :onos onobræ li rep elatibro ammargaid li e inorttele id enoizaruignoc aL .JiluaP id enoisulsee id oipincirp li noc odrocca ni(LM octinaug oremun orol len onocsirefid SM e L .N icitnaug iremun onnah p2 onobræ id ilatibro ilgen inorttele eud i ,otnatreP ,itaiippocca non inorttele id omisssam oremun li ereva eht :Â itareneged ilatibro id eires anu id onrethi'lla inorttele noc omota nu rep assab 'Aip aigreine id enoizaruignoc al .dnuH id aloger allad ottricsed emoc itipmeir onas ilatibro ilG .Itareneged am ,isrevid ilatibro eud ni itaipocca non inorttele ilg eraiscal id o inorttele ilg eranibba id e p2 ilatibro ilged onu eripmeir id atlecs al omaibba arO ,p2 llehsottos al onapucco inorttele eud itnatser I .S2 e S1 ilatibro ilg onoipeir orol id orttauQ .inorttele ies ah j6 ocimota oremun(onobræ li .odneiimeir omaits eht ottehcstos assets allen itouu ilatibro ilautneve erataneserppar rep etouu elotacs omaidulcni ,ilatibro immargaid omaignesis odnauQ .ilatibro itseuq id onu erapucco 'Aup enorttele' e)1+ ,0 ,1âEË = lm(itareneged p2 ilatibro ert onotiseE .p2 elatibro nu .Aras eht ,aigreine id olleuil omisssorp li erapucco eved enorttele otiaug li ,inorttele eud olos erenentoc 'Aup S llehsottos isaislaug ©AhcioP .2 = n oicsug li onnarepucco inorttele ert e inorttele eud noc otipmeir 'Â 1 = n oicsug li .inorttele euqnic eneitnoc j5 ocimota oremun(orob id omota nU .S2 elatibro'llen etnenamir oizaps ol eipmeir enorttele orttauQ li .oelcun li onoadnoric eht inorttele orttauQ e oelcun len inotorp orttauQ eneitnoc .4 id ocimota oremun nu noc ,onilacla arret id ollatem id oilireb led omota nu :onos oitli led elatibro ammargaid e enoizaruignoc p2 p2 nu olos ah j9 ocimota oremun(orouff li ,eud irtla ilged onucaic ni enorttele olognis nu e)itsoppo irig onnah containing an unrelated electron. All electrons in noble neon gas (atomic number 10) are coupled, and all orbitals in the $n = 1$ and $n = 2$ shells are filled. The electrons configurations and the orbital diagrams of these four elements are: alkali metal sodium (atomic number 11) has an electron in more than the neon atom. This electron must enter the more low-energy subshell available, the orbital 3s, giving a 1S22S2P63S1 configuration. The electrons that occupy the most orbital external shell (l) (the highest value of N) are called Valenza electrons, and those who occupy the orbital of the internal shell are called nucleus electrons (figure (PageIndex{5})). Since The nucleus electrons coverage correspond to electron configurations of noble gas, we can shorten the electrons configurations by writing the noble gas that corresponds to the configuration of the basic electron, together with the electrons of value in a condensed format. For our sodium example, the [ne] symbol represents Core electrons, (1s22s22p6) and our abbreviated or condensed configuration is [ne] 3s1. With the noble gas symbol whose configuration corresponds to the configuration of the electron of the nucleus of the other element. In the same way, the abbreviated configuration of the lithium can and SSSre represented as [He] 2S1, where [he] represents the configuration of the atom of helium, which is identical to that of the internal shell filled with lithium. Write the configurations in this way emphasizes the similarity of lithium and sodium configurations. Both atoms, which are in the family of alkali metals, have only an electron in a background of value outside with internal shells. See: the magnesium alkaline atipmeir atipmeir s anu onnah imota ilg ibmartne .2s2]eH(.aigimaf id oilireb ou la ogolana 'Â ,2s3]eN(enoizaruignoc anu ni inorttele 21 ious i noc),21 ocimota oremun(arret outside their inner covers filled. Aluminium (atomic number 13), with 13 electrons and electron configuration [Ne]3s23p1, is analogous to the family boron, [He]2s22p1. The configurations of silicon electrons (14 electrons), phosphorus (15 electrons), sulphur (16 electrons), chlorine (17 electrons), and argon (18 electrons) are analogous in the electron configurations of their outer shells to the corresponding members of the carbon family, nitrogen, oxygen, fluorine and neon, respectively, except that the number of main quantum of the outer shell of the heaviest elements is increased from one to zero = 3. Figure \ (PageIndex{6}): This version of the periodic table shows the external-shell electron configuration of each element. Note that each group, configuration is often similar. When we arrive at the next element in the periodic table, alkali metal potassium (atomic number 19), we expect that we will begin to add electrons to subshell 3d. However, all available chemical and physical tests indicate that potassium is similar to lithium and sodium, and that the next electron is not added to level 3d but is, instead, added to level 4s (Figure \ (PageIndex{3}) or \ (PageIndex{4})). As discussed earlier, the 3d orbit without radial nodes is higher in energy because it is less penetrating and more shielded by the nucleus than the 4, which has three radial nodes. Thus, potassium has an electron configuration of [Ar]4s1. So, potassium corresponds to Li and Na in its configuring valence shell. The next electron is added to complete the 4s subshell and the calcium has an electron configuration of [Ar]4s2. This gives calcium an external-shell electron configuration corresponding to that of beryllium and magnesium. 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